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Molarity Practice I 11. How many grams of potassium chloride are needed to make 200 ml, of a 2.5 M solution?

(Answer:  $2.5\text{M} = 0.5\text{mol/L}$ )

12. What is the molarity of a 5.00 x 10<sup>2</sup> ml, solution containing 249 g of calcium iodide? (Answer: 1.69 M)

13. 0.5 L

**Solutions and Molarity  
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Read Book Molarity Worksheet Answers Page 68 Molarity =  $58.5 \text{ g} = 1.69 \text{ M}$   $0.562 \text{ L}$  .  
 $10.78.9 \text{ g}$  of  $\text{CuSO}_4 \cdot 8\text{H}_2\text{O}$  is dissolved in  $500.0 \text{ mL}$  of water, calculate the

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68 molarity.  $78.9 \text{ g} \times 1 \text{ mole}$  .  
Molarity =  $303.76 \text{ g} = 0.519$   
M  $0.5000 \text{ L}$  . Stoichiometry  
Worksheet # 3 Molarity  
Worksheet # 1 - W.J. Mouat  
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Program What is the molarity

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of the following solutions given that: 1) 1.0 moles of potassium fluoride is dissolved to make 0.10 L of solution. 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution.

## **Molarity Worksheet W 331 - Everett Community College**

Molarity Worksheet # 2  
identifiera \_\_\_\_\_ What does molarity mean? Number of moles of solute. 1 liter solution. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution?  $4.53 \text{ mol LiNO}_3 = 1.59 \text{ M LiNO}_3$ . 2.85 L soln

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## **Molarity Worksheet 2 ANSWERS - Google Docs**

Molarity = \_\_\_\_\_ Problems:  
Show all work and circle  
your final answer. 1. To  
make a 4.00 M solution, how  
many moles of solute will be  
needed if 12.0 liters of  
solution are required? 2.  
How many moles of sucrose  
are dissolved in 250 mL of  
solution if the solution  
concentration is 0.150 M? 3.  
What is the molarity of a  
solution of HNO

### **Worksheet: Molarity Name**

Molarity+calculations+(fillN  
inalltheboxes)+  
++solute+molesof+ solute+  
grams+of+ solute+ volumeof++

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68 solution+ Concentration+  
(Molarity, +M=mole/L)+  
++NaCl+

## **Molarity Molality Osmolality Osmolarity Worksheet and Key**

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Chemistry If8766 Molarity  
Answers Molarity Chemistry  
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solid.braziljs.org The  
answer is 1.00 mol/L. Notice  
that both the units of mol  
and L remain. Neither  
cancels. A symbol for mol/L  
is often used.

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Worksheet: Molarity ANSWER  
KEY PART 1 Molarity: a quantitative description of solution concentration. Abbreviated M Molarity = moles of solute  $\div$  liters of solution Problems: Show all work and circle your final answer. 1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of

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68 solution are required? moles  
= Molarity x liters = 4.00 M  
x 12.0 L = 48 moles of  
solute

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