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been taught about the
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dividing numbers, you
need to round your
answers to the correct
number of significant
figures. To round

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correctly, follow these simple steps: 1)

Count the number of significant figures in each number. 2)

Round your answer to the least number of significant figures.

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of Atoms. (Click on the previous link to download a copy.) At this point, the students have not been taught about the structure of an atom. They are not familiar with terms such as atomic number, mass number or isotopes. Use the Guided

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53 concept of

significant figures has

the answer. When

multiplying and

dividing numbers, you

need to round your

answers to the correct

number of significant

figures. To round

correctly, follow these

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simple steps: 1)

Count the number of significant figures in each number. 2)

Round your answer to the least number of

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X = chemical symbol of the element Z =
“atomic number” A =
“mass number” , ,

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and are notations that represent isotopes of carbon. ${}^1_1\text{H}$, and are notations that represent isotopes of hydrogen. The part of the atom where the protons and neutrons are is called the nucleus. ${}^{12}_6\text{C}$ ${}^{13}_6\text{C}$ ${}^{14}_6\text{C}$ ${}^{12}_6\text{C}$ ${}^{13}_6\text{C}$ ${}^{14}_6\text{C}$ ${}^1_1\text{H}$ ${}^1_1\text{H}$ ${}^2_1\text{H}$ ${}^1_1\text{H}$ ${}^3_1\text{H}$ ${}^3_1\text{H}$

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During a chemical reaction, new substances are formed. Reactants are transformed into different products.

Atoms are never created or destroyed, but they are rearranged. A chemical equation

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represents what

happens during a
reaction. The

following is an

example of a

chemical equation:

Example Equation:

$\text{Ca} + \text{HNO}_3$ (

$\text{Ca}(\text{NO}_3)_2 + \text{H}_2$

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Hint: add your answer to question 1 and four times your answer to question 2. $4 + 4(7) = 4 + 28 = 32$

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Burners, test-tubes
filled with brightly

coloured fluids,

goggles and strange

smells; this is the

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least to someone

starting high school!

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practical subject that

is at the heart of our

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