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(03.12.2017) **Chemical Equilibrium -02**

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One thing to note about equilibrium is that the reactions do not stop; both the forward reaction and the reverse reaction continue to occur. They both occur at the same rate, so any overall change by one reaction is cancelled by the reverse reaction. We say that chemical equilibrium is dynamic, rather than static. Also, because both

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reactions are occurring simultaneously, the equilibrium can be written backward.

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Equilibrium - Study.com

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now, chemical systems have been treated
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$\text{MgF}_2(\text{s}) \rightleftharpoons \text{Mg}^{2+}(\text{aq}) + 2 \text{F}^{-}(\text{aq})$ In a saturated solution of MgF_2 at 18°C , the concentration of Mg^{2+} is 1.21×10^{-3} molar. The equilibrium is represented by

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the equation above. (a) Write the expression for the solubility-product constant, K_{sp} , and calculate its value at 18°C .

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nearly every reaction that we will study in this course. Many reactions seek to achieve equilibrium and this unit represents the first look that we will take at how to characterize aqueous and gaseous equilibrium systems through a wide variety of chemical reactions.

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Questions that look like the questions you will see on test day. Atomic Structure and Properties

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study of matter, its composition and the

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acidic/basic salts

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In a chemical reaction, chemical
equilibrium is the state in which the

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forward reaction rate and the reverse reaction rate are equal. The result of this equilibrium is that the concentrations of the reactants and the products do not change.

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chemical equilibrium? A state of balance
in a reaction where the forward and
reverse reaction speed is unequal. A state
where the reaction is unbalanced. Page 5/7

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reaction: Equilibrium is defined as the

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state in which the rate of the forward reaction and the rate of the reverse reaction is equal. Hence, in this state, the...

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